

## Weak Electrolytes

- **Electrolytes** are substances that dissolve in water and acquire the capacity to conduct electricity.

### WEAK ELECTROLYTES

- A weak electrolyte is an electrolyte that **doesn't dissolve completely** in water. It dissolves only 1-10% in water.
- **Weak acids and weak bases** make a weak electrolyte.
- An example of a weak electrolyte is **acetic acid**.
- When acetic acid dissolves in water most of its original molecule remains as it is. A few of them get converted into ions, making this a weak electrolyte.
- $\text{CH}_3\text{COOH} + \text{H}_2\text{O} \rightleftharpoons \text{CH}_3\text{COO}^- + \text{H}_3\text{O}^+$

## Strong Electrolytes

SOURCE

- A **strong electrolyte** is an electrolyte that dissolves almost completely in water.
- **Strong acids and strong bases** make strong electrolytes.
- An example of a strong electrolyte is **Hydrogen Chloride**.
- Hydrogen chloride in its pure molecular state is a nonelectrolyte. But when it's dissolved in water, the HCl molecules get converted into  $\text{H}^+$  and  $\text{Cl}^-$  ions.

